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DETERMINATION OF THE NUMBER OF ELECTRONS IN THE
ELECTROOXIDATION OF DIETHYLAMINO-4-METHYL-HEXIN-2-OLA-4

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When studying the mechanism of the electrode oxidation process of diethylamino-4-methyl-hexine-2-ol-4 (DEAMGO), first of all, it was necessary to determine the number of electrons given up during its electroelectrooxidation.

The number of electrons involved in the oxidation of one depolarizer molecule on a platinum rotating microdisk electrode could be determined on the basis of the theoretical equation of the mathematical description of the process for the limiting current on such an electrode:

$$I = 0,62 n F S D^{2/3} V G C$$

where, n is the number of electrons; F -Faraday constant, C/mol^{-1} ; S is the surface area of the platinum disc, cm^2 ; D -diffusion coefficient

depolarizer, cm^2/s ; V -kinematic viscosity of the background solution, cm^2/s ; G angular speed, rev/s ; C -concentration of the depolarizer, mol/l , (mol/cm^3); I value of current, A .

Note, for this purpose, the approximate value of the diffusion coefficient of the reagent used, calculated by the Stokes-Einstein formula, is not very reliable, in the capacity for many-electron processes.

In this regard, it was decided to determine the number of electrons involved in the electrooxidation of one DEAMHO molecule in the studied non-aqueous solutions in the presence of supporting electrolytes of various nature by the coulometric method, where the number of electrons is calculated directly from the Faraday formula:

$$n = \frac{Q \cdot M}{F \cdot P}$$

The mass of the electrooxidized reagent was found by the difference between the taken and the remaining amount and after the study using the formula:

$$F = Co \cdot Vo - Ck \cdot Vk$$



The amount of electricity was found by graphical integration of the current value corrected for the background current, ranging from zero to the time of the study, that is, by determining the area of the figure limited between the curves of the dependence of the current value on time during the electrolysis of the DEAMGO reagent solution, electrolysis of the background and the ordinates $t = 0$ and $t = t_{el}$. : (tel-time of electrolysis).

Denoting the area of the figure obtained in this way through S_1 , we have the formula:

$$Q = S_1 \cdot a \cdot b$$

Putting the expressions for P and Q into the equation, we get the final formula for determining the number of electrons:

$$n = \frac{S_1 \cdot a \cdot b \cdot M}{F \cdot (C_o \cdot V_o) - (C_k \cdot V_k)}$$

The values of S_1 , a , b , C_o , C_k , V_o , V_k and the numbers of electrons found from them for the investigated reagent - DEAMHO in anhydrous acetic acid, npropanol, DMF (DMSO).

As can be seen from the calculation data, during the electrooxidation of DEAMGO on a planned disk microanode in the studied protolytic media, the number of electron donation is close to 2, which confirmed our assumption.

LITERATURE

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